

## Parts of the Atom

*\*Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit.)*

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	MASS (in AMU's)
proton			
neutron			
electron			

*\*Here's how to predict the structure of any element on the periodic table. For example, look at carbon:*

<p>6 C carbon 12.011</p>
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←atomic number: This is the number of protons (also the number of electrons)

←chemical symbol: This is an abbreviation for the element.

←element name: This is the name of the element.

←atomic mass (in AMU's): This number (rounded off) tells you the number of protons *plus* the number of neutrons. (Electrons are too small to be included.)

**To find the number of protons, look at the atomic number:**

Example for carbon: Atomic number: 6      Number of protons: 6

**To find the number of neutrons, subtract the atomic number from the atomic mass:**

Example for carbon: Atomic mass: 12 (rounded)      Number of neutrons:  $12 - 6 = 6$

**To find the number of electrons, look at the atomic number:**

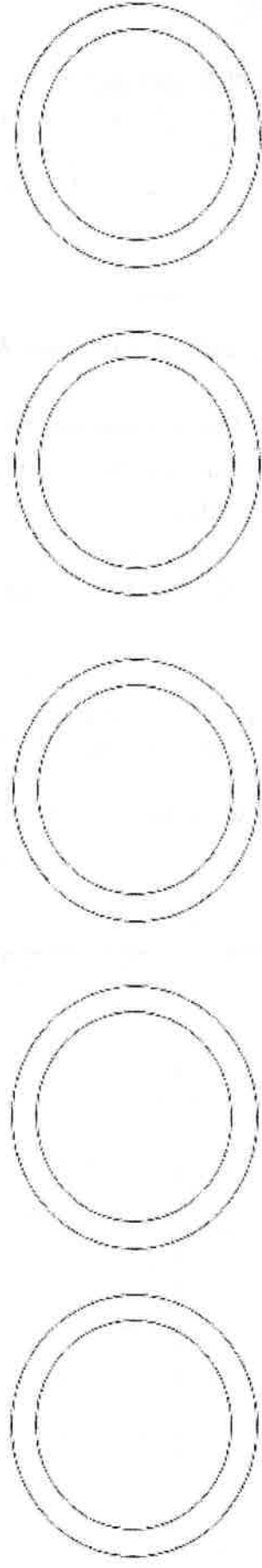
Example for carbon: Atomic number: 6      Number of electrons: 6

*\*Complete the table for the elements with atomic numbers 1-10.*

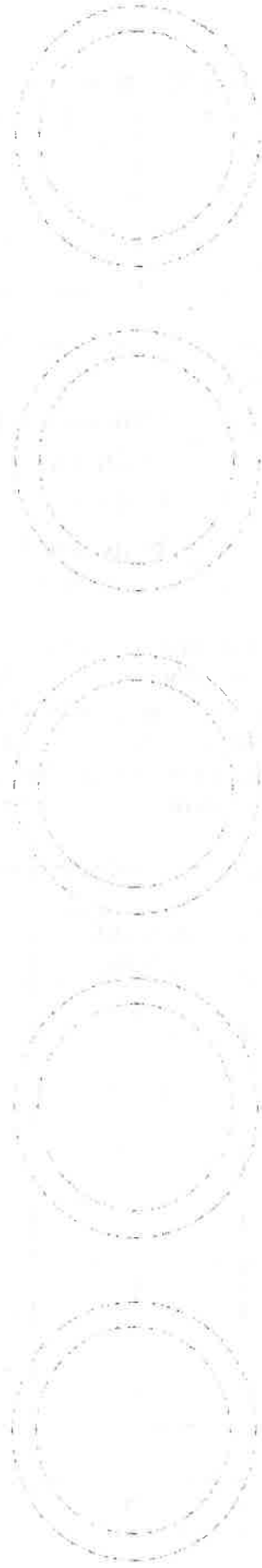
ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
	1				
	2				
	3				
	4				
	5				
	6				
	7				
	8				
	9				
	10				

Use the spaces below to draw the correct amount of protons, neutrons, and electrons for the first 10 elements. Use your chart and the periodic table to help. BE SURE to fill in the key to the right. Use a different color for each of the three main atomic parts.

O = Protons
O = Neutrons
o = Electrons



1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_



6. \_\_\_\_\_
7. \_\_\_\_\_
8. \_\_\_\_\_
9. \_\_\_\_\_
10. \_\_\_\_\_